

**Melting Point**

660°C

Boiling Point

2467°C

Density2.7 g/cm³**Appearance****Other Physical Properties**

Aluminum is a good conductor of heat.

Chemical Properties

Aluminum reacts quickly with oxygen in the air to form a protective layer of oxide.

Compounds

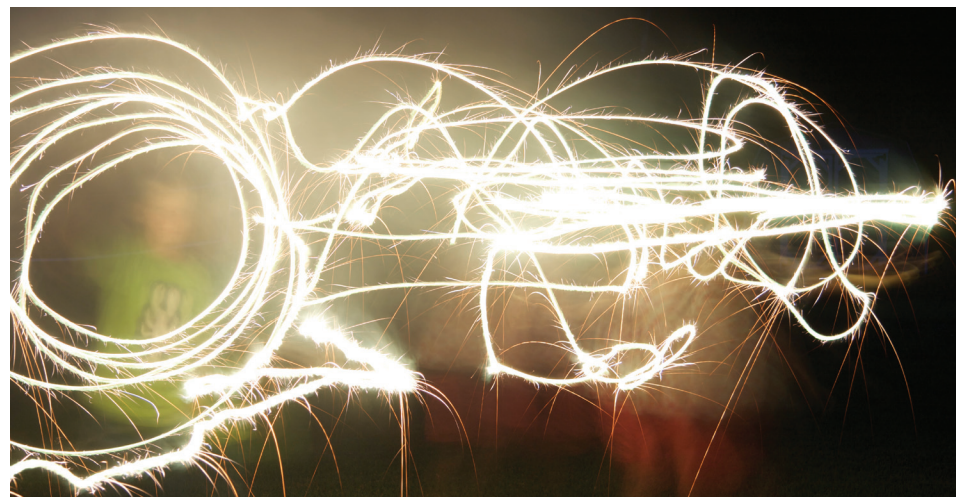
Compounds of aluminum are white in color.

Uses

Aluminum is used in cookware, foil, aircraft, overhead cables, and window frames.

Notes

Aluminum is obtained by the electrolysis of its molten ore, alumina (aluminum oxide). It is the third most abundant element in Earth's crust.



IT IS ONLY A COATING OF ALUMINUM OXIDE THAT MAKES ALUMINUM SEEM UNREACTIVE. IN FACT, ALUMINUM IS A REACTIVE METAL, AND THIS PROPERTY IS EXPLOITED TO MAKE SPARKLERS. ALUMINUM IS ALSO USED AS FUEL IN THE SOLID FUEL BOOSTERS OF NASA'S SPACE SHUTTLE.

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ALUMINUM HAS A LOW DENSITY. THIS PROPERTY AND ITS RESISTANCE TO CORROSION MAKE IT AN IDEAL MATERIAL FOR AIRCRAFT MANUFACTURE.

PHOTO: National Air and Space Museum, Smithsonian Institution (SI 2003-29162)