

**Melting Point**

3550°C

**Boiling Point**

4827°C

**Density**1.9 g/cm<sup>3</sup> (charcoal); 2.3 g/cm<sup>3</sup> (graphite, or carbon rod); 3.5 g/cm<sup>3</sup> (diamond)**Appearance****Other Physical Properties**

There are three forms of carbon: charcoal, graphite, and diamond (which is the hardest substance known). Each form has different physical properties.

**Chemical Properties**

Carbon combines with oxygen in the air to form carbon dioxide.

**Compounds**

Carbon forms a wide variety of compounds, many of which are important components in living things. It is present in the compounds found in coal and limestone.

**Uses**

Charcoal is used in filters to absorb impurities in water or smells from the air. It is also used as a fuel for cooking (for example, in barbecues). Graphite is used as a lubricant and to make pencil leads. Diamond is used in jewelry and in cutting tools (because of its hardness).

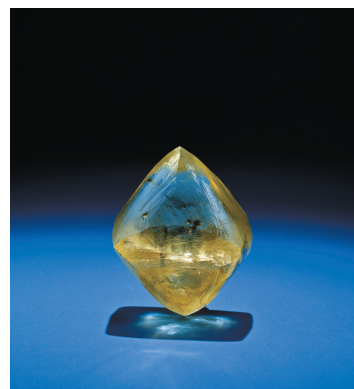
**Notes**

Carbon is the sixth most abundant element in the universe. Combined with other elements, carbon is found in all substances made from oil, including plastics.



▶ THESE DIAMONDS HAVE BEEN CUT TO MAKE JEWELRY.

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▶ THE OPPENHEIMER DIAMOND IS ONE OF THE LARGEST UNCUT DIAMONDS IN THE WORLD.

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